

Mg(2) and in addition by K(2). Atom O(3) appears therefore to be locked in, which is reflected in the fact that the amplitude of its thermal motion is the smallest of the three oxygen atoms and that its motion seems to be isotropic within the limits of error. The details of the orientation of the thermal ellipsoids thus are in qualitative agreement with the geometrical arrangements in this crystal structure.

Table 7. $K_2Mg_5Si_{12}O_{30}$, root-mean-square thermal displacements along principal axes and their orientations relative to *a*, *b* and *c*

	Axis	Displacement	Angle with axis		
			[100]	[010]	[001]
K(1)	1	0.111 (4) Å	—	—	90°
	2	0.111	—	—	90
	3	0.154 (5)	90°	90°	0
K(2)	1	0.106 (5)	—	—	90
	2	0.106	—	—	90
	3	0.136 (6)	90	90	0
Mg(1)	1	0.077 (5)	—	—	90
	2	0.077	—	—	90
	3	0.119 (5)	90	90	0
Mg(2)	1	0.087 (6)	90	30	90
	2	0.112 (4)	90	90	0
	3	0.116 (5)	0	120	90
Si	1	0.088 (2)	68 (34)	87 (16)	28 (29)
	2	0.090 (2)	157 (32)	64 (5)	67 (33)
	3	0.109 (2)	98 (5)	26 (5)	104 (5)
O(1)	1	0.095 (6)	90	90	0
	2	0.155 (6)	32 (85)	152 (85)	90
	3	0.158 (5)	58 (85)	62 (85)	90
O(2)	1	0.105 (5)	27 (5)	145 (4)	81 (7)
	2	0.132 (4)	78 (8)	91 (8)	166 (8)
	3	0.157 (4)	67 (4)	55 (4)	79 (8)
O(3)	1	0.108 (5)	99 (32)	21 (29)	96 (21)
	2	0.113 (5)	158 (27)	73 (34)	109 (24)
	3	0.121 (4)	110 (24)	78 (17)	20 (23)

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The Crystal Structures of Polymorphs I and III of Sulphathiazole

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The crystal structure of sulphathiazole II is known. The crystal structures of the other two known crystalline forms of sulphathiazole (*N'*-2-thiazolylsulphanilamide) have now been determined and are reported. The intensities were measured using Mo *K* α radiation and an automatic diffractometer. The structures were solved by the application of the Karle-Hauptman Σ_2 relationship and refinement was done by full-matrix least-squares methods. The main differences between the polymorphs of sulphathiazole lie in the hydrogen bonding and its effects on the arrangement of the molecules in the crystals.

Introduction

In the paper reporting the structure determination of sulphathiazole II (Kruger & Gafner, 1972) crystallographic data for the other two known polymorphs,

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modifications I and III, were listed. This information is repeated here in Table 1. In order to investigate the structural differences involved in the polymorphism of sulphathiazole, the crystal structures of sulphathiazole I and III have now been determined.

Experimental

The intensity data were collected from roughly cubical crystals with side lengths not exceeding 0.2 mm using a Hilger & Watts four-circle automatic diffractometer. In both cases the intensities were measured using the ω -scan method and Zr-filtered Mo radiation. Only the unique reflexions with $2\theta \leq 40^\circ$ were measured for sulphathiazole I since the proportion of unobservable intensities became too high above this limit. All reflexions with $2\theta \leq 55^\circ$ were measured for sulphathiazole III. Corrections for absorption were not applied as μR was less than 0.05 in both cases.

Determination of the structures

It was not found possible to use the heavy-atom method owing to the presence of two molecules in the asymmetric unit of each crystal. The structures were

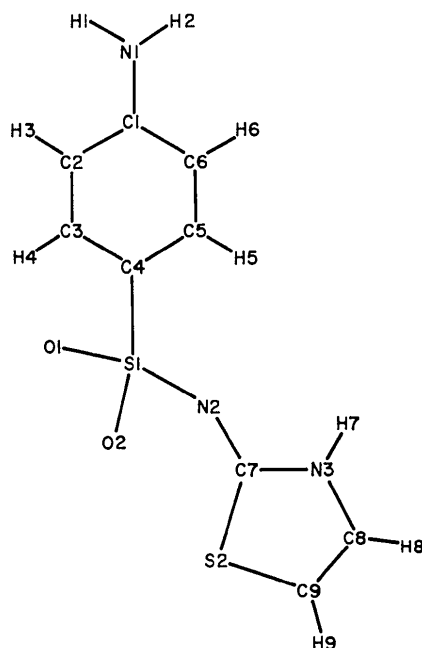


Fig. 1. The numbering of atoms in the sulphathiazole molecule.

thus solved by the symbolic addition procedure using the phase determination programs included in the *X-ray 67* (Stewart, 1967) computing system.

In the case of sulphathiazole I, the phases of 173 reflexions with E values greater than 1.5 were determined and the E map calculated from them revealed the positions of the four sulphur atoms in the asymmetric unit. A Fourier synthesis phased on the contributions of the sulphur atoms gave the positions of the other 28 non-hydrogen atoms.

All reflexions with E values greater than 1.4 were used for the phase determination of sulphathiazole III. The phase determination process gave two possible sets of phases for the 362 reflexions. E maps were calculated with both sets. Only one of these could be interpreted and this gave the positions of the 32 non-hydrogen atoms.

The refinement of atomic parameters

Refinement was done by full-matrix least-squares procedures using the program *ORFLS* of Busing, Martin & Levy (1962). The atomic scattering factors used were those given by Hanson, Herman, Lea & Skillman (1964). Only the observed reflexions were used in the refinement of the structures (Dunning & Vand, 1969), that is, all the intensities higher than $2\sigma(I)$ where $\sigma(I)$ is the estimated standard deviation in intensity as determined from counting statistics.

(a) Sulphathiazole I

Of the 3328 reflexions recorded, 1843 were used in the refinement. After refinement of the positional and anisotropic thermal parameters, a three-dimensional difference Fourier synthesis was calculated. All hydrogen atoms could be found with the exception of the four bonded to amine nitrogen atoms. In subsequent refinement the 14 hydrogen atoms which could be located were included and given the anisotropic thermal parameters of the atoms to which they were bonded. Only their positional parameters were varied.

The most satisfactory method of weighting was found to be a modification of the weighting scheme proposed by Hughes (1941), where

Table 1. *Crystal data for the polymorphs of sulphathiazole*

Polymorph	I	II	III
Melting point ($^\circ\text{C}$)	200–202	200–202	173–175 200–202 or 173–175
Transition point ($^\circ\text{C}$)	–	173–175	–
Habit	elongated rods	hexagonal prisms	hexagonal plates
Space group	$P2_1/c$	$P2_1/c$	$P2_1/c$
a (\AA)	10.554 (5)	8.235 (4)	17.570 (9)
b (\AA)	13.220 (7)	8.550 (4)	8.574 (4)
c (\AA)	17.050 (8)	15.558 (8)	15.583 (8)
β ($^\circ$)	108.06 (1)	93.67 (1)	112.93 (1)
Z	8	4	8
D_m (g.cm^{-3})	1.50	1.55	1.57
D_c (g.cm^{-3})	1.499	1.550	1.567

Table 2 (cont.)

A large table of numerical data with multiple columns, likely representing coordinates or statistical values for different samples. The table is organized into columns and rows, with some columns containing headers like 'K, L'.

angles in the four molecules are listed in Table 5. The big differences between them are the only significant quantitative differences between the molecular geometry in the polymorphs of sulphathiazole.

The crystal structures of sulphathiazole I and III are illustrated by the stereo pairs (Johnson, 1965) of their unit-cell contents in Figs. 2 and 3. Hydrogen bonds

distances and other close intermolecular approaches are listed in Table 6.

In sulphathiazole I, centrosymmetric dimers are formed by pairs of molecules through symmetrical hydrogen bonds between nitrogen atoms N(2) and N(3). Separate dimers are bonded together through hydrogen bonds between atoms N(1) and O(1) and O(2). Only

Table 3. *Positional and thermal parameters*(a) Atomic parameters ($\times 10^4$) with e.s.d.'s for the non-hydrogen atoms of sulphathiazole I and III.

Positional parameters are given as fractional coordinates and thermal parameters are given according to the expression

$$T = \exp [-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + 2\beta_{12}hk + 2\beta_{13}hl + 2\beta_{23}kl)] .$$

The e.s.d.'s are given in brackets and refer to the least significant figures in the parameter values. The values given for each parameter are from top to bottom, for sulphathiazole I, molecules *A* and *B*, and sulphathiazole III, molecules *A* and *B* respectively.

	<i>x</i>	<i>y</i>	<i>z</i>	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
S(1)	3221 (3)	709 (2)	1220 (2)	109 (3)	71 (2)	40 (1)	- 1 (2)	11 (2)	9 (1)
	8063 (2)	4260 (2)	884 (2)	84 (3)	84 (2)	32 (1)	- 2 (2)	10 (1)	11 (1)
	3829 (1)	3982 (2)	5584 (1)	22 (1)	87 (2)	29 (1)	3 (1)	5 (1)	1 (1)
S(2)	8916 (1)	3802 (2)	8178 (1)	23 (1)	77 (2)	27 (1)	4 (1)	3 (1)	0 (1)
	1301 (3)	4338 (3)	4303 (2)	90 (3)	149 (3)	57 (2)	- 14 (3)	- 9 (2)	25 (2)
	6434 (3)	2537 (2)	4660 (2)	157 (4)	61 (2)	72 (2)	- 3 (3)	36 (2)	6 (2)
O(1)	3558 (1)	1741 (2)	3781 (1)	64 (1)	139 (3)	29 (1)	16 (1)	13 (1)	1 (1)
	8635 (1)	1513 (2)	6404 (1)	71 (1)	128 (3)	26 (1)	13 (1)	12 (1)	1 (1)
	1868 (6)	402 (5)	1801 (4)	106 (8)	96 (6)	68 (4)	- 23 (5)	29 (4)	10 (4)
O(2)	8375 (6)	5270 (5)	1193 (4)	146 (9)	79 (6)	60 (4)	- 32 (6)	2 (4)	13 (4)
	4314 (2)	4781 (4)	6445 (2)	28 (2)	100 (6)	41 (2)	3 (3)	- 1 (2)	- 10 (3)
	9397 (2)	4611 (4)	9027 (3)	33 (2)	102 (6)	42 (2)	- 3 (3)	- 1 (2)	- 20 (3)
N(1)	4206 (7)	345 (6)	1951 (4)	167 (10)	90 (6)	48 (4)	12 (7)	7 (5)	20 (4)
	8726 (6)	3875 (6)	342 (4)	120 (9)	162 (8)	36 (3)	12 (7)	29 (4)	18 (4)
	3761 (2)	4755 (5)	4733 (3)	31 (2)	141 (6)	44 (2)	6 (3)	14 (2)	27 (3)
N(2)	8870 (2)	4502 (5)	7323 (3)	33 (2)	136 (7)	40 (2)	3 (3)	13 (2)	20 (3)
	3359 (8)	5164 (7)	1483 (5)	197 (13)	53 (7)	45 (4)	- 1 (8)	26 (6)	7 (4)
	9410 (8)	1607 (7)	3827 (5)	169 (12)	93 (8)	49 (4)	22 (8)	26 (6)	38 (5)
N(3)	449 (3)	3047 (7)	5498 (3)	26 (3)	95 (7)	37 (3)	- 4 (7)	7 (3)	1 (3)
	5538 (3)	3230 (6)	8172 (3)	31 (2)	104 (8)	43 (3)	7 (4)	10 (2)	13 (4)
	3758 (7)	345 (6)	476 (5)	87 (8)	77 (6)	41 (4)	6 (6)	8 (5)	- 1 (4)
C(1)	6470 (7)	4295 (6)	462 (4)	106 (9)	66 (6)	27 (3)	6 (7)	10 (4)	- 4 (4)
	4202 (3)	2246 (5)	5693 (3)	30 (2)	108 (8)	29 (2)	19 (3)	7 (2)	2 (4)
	9258 (3)	2050 (5)	8318 (3)	29 (2)	82 (6)	29 (2)	11 (3)	4 (2)	- 1 (3)
C(2)	3433 (6)	4937 (6)	4086 (5)	88 (9)	66 (6)	36 (4)	- 10 (6)	- 7 (5)	1 (4)
	4550 (8)	1308 (6)	4508 (5)	109 (10)	58 (6)	47 (4)	12 (7)	4 (5)	7 (4)
	4107 (3)	- 219 (5)	5081 (3)	33 (2)	102 (7)	40 (3)	0 (3)	15 (2)	- 2 (4)
C(3)	9202 (3)	- 439 (5)	7706 (3)	29 (2)	83 (7)	33 (2)	4 (3)	14 (2)	- 4 (3)
	3324 (9)	4137 (8)	1383 (5)	104 (13)	83 (9)	22 (4)	- 4 (9)	14 (5)	5 (5)
	9085 (9)	2234 (8)	3144 (6)	78 (11)	69 (9)	50 (5)	- 2 (8)	16 (6)	- 1 (5)
C(4)	1233 (3)	3298 (6)	5491 (3)	23 (2)	79 (8)	32 (3)	- 3 (3)	9 (2)	- 10 (4)
	6334 (3)	3374 (6)	8142 (3)	28 (2)	75 (7)	28 (3)	2 (3)	5 (2)	1 (4)
	4467 (9)	3558 (8)	1703 (6)	103 (12)	66 (8)	49 (5)	- 8 (9)	11 (6)	- 8 (5)
C(5)	9512 (9)	1992 (7)	2456 (6)	151 (14)	66 (7)	44 (5)	38 (9)	30 (7)	9 (5)
	1357 (3)	4184 (6)	4813 (3)	25 (2)	81 (8)	24 (3)	4 (3)	2 (2)	- 4 (4)
	6452 (3)	4135 (6)	7420 (3)	20 (2)	99 (9)	28 (3)	8 (4)	1 (2)	2 (4)
C(6)	4466 (8)	2528 (10)	1644 (5)	56 (9)	102 (10)	38 (5)	8 (9)	- 10 (5)	1 (6)
	9145 (9)	2582 (9)	1766 (6)	130 (12)	93 (9)	35 (5)	20 (10)	28 (6)	- 1 (6)
	2152 (3)	4391 (6)	4831 (3)	27 (2)	100 (8)	25 (2)	1 (3)	7 (2)	- 2 (4)
C(7)	7225 (3)	4267 (6)	7408 (3)	33 (3)	83 (8)	23 (2)	6 (4)	6 (2)	6 (3)
	3281 (8)	2031 (7)	1254 (5)	83 (10)	66 (7)	27 (4)	12 (7)	9 (5)	15 (4)
	8400 (8)	3456 (7)	1738 (5)	73 (11)	63 (7)	33 (4)	3 (8)	5 (5)	- 7 (4)
C(8)	2830 (3)	3743 (6)	5548 (3)	23 (2)	76 (7)	23 (2)	2 (3)	4 (2)	- 4 (4)
	7900 (3)	3630 (6)	8141 (3)	25 (2)	76 (7)	28 (2)	- 2 (3)	7 (2)	- 2 (3)
	2136 (9)	2591 (9)	908 (6)	90 (11)	83 (8)	44 (5)	- 12 (9)	6 (6)	- 9 (6)
C(9)	8023 (8)	3711 (7)	2416 (5)	73 (10)	66 (7)	36 (4)	5 (7)	16 (5)	- 8 (5)
	2710 (4)	2866 (7)	6237 (4)	25 (3)	109 (8)	25 (3)	- 2 (4)	2 (2)	3 (4)
	7778 (4)	2888 (7)	8867 (4)	33 (3)	107 (8)	27 (3)	19 (4)	4 (2)	11 (4)
C(10)	2152 (8)	3628 (9)	974 (5)	68 (11)	91 (9)	44 (4)	8 (9)	- 7 (5)	- 7 (5)
	8335 (9)	3114 (7)	3095 (5)	96 (11)	78 (8)	28 (4)	17 (8)	19 (5)	14 (4)
	1943 (4)	2670 (6)	6225 (4)	32 (3)	80 (9)	33 (3)	- 3 (4)	12 (2)	9 (4)
C(11)	6991 (4)	2770 (6)	8880 (4)	30 (3)	103 (8)	26 (3)	8 (4)	9 (2)	13 (4)
	2970 (9)	4654 (7)	4707 (6)	110 (11)	44 (7)	48 (5)	13 (7)	11 (6)	1 (5)
	5873 (9)	1402 (8)	4944 (6)	116 (13)	63 (7)	38 (4)	- 9 (8)	30 (6)	- 17 (5)
C(12)	3982 (3)	1324 (7)	4958 (4)	28 (2)	112 (9)	37 (3)	7 (4)	14 (2)	12 (4)
	9062 (3)	1101 (7)	7600 (3)	18 (2)	103 (8)	30 (3)	2 (4)	2 (2)	1 (4)
	2544 (11)	4866 (8)	3319 (6)	138 (15)	91 (9)	45 (5)	- 26 (10)	- 9 (7)	1 (5)
C(13)	4053 (12)	2068 (10)	3951 (7)	159 (18)	74 (8)	68 (7)	38 (11)	12 (9)	7 (6)
	3875 (4)	- 1084 (7)	4268 (4)	46 (3)	126 (9)	49 (3)	1 (5)	16 (3)	23 (5)
	8972 (4)	- 1309 (7)	6907 (4)	35 (3)	139 (10)	47 (4)	5 (5)	16 (3)	- 20 (5)
C(14)	1388 (11)	4521 (11)	3310 (7)	129 (14)	148 (13)	60 (6)	- 51 (12)	- 23 (7)	24 (7)
	4924 (14)	2791 (9)	3957 (7)	262 (22)	47 (9)	76 (7)	10 (11)	44 (10)	13 (6)
	3562 (5)	- 217 (8)	3504 (4)	58 (4)	156 (11)	39 (3)	9 (6)	10 (3)	- 13 (5)
C(15)	8648 (5)	- 452 (7)	6141 (4)	75 (5)	123 (10)	47 (4)	15 (6)	26 (4)	- 21 (5)

Table 3 (cont.)

(b) Fractional coordinates ($\times 10^3$) with e.s.d.'s for the hydrogen atoms of sulphathiazole I and III

	<i>x</i>	<i>y</i>	<i>z</i>
H(1)	—	—	—
	22 (5)	327 (10)	494 (6)
	508 (4)	356 (9)	770 (5)
H(2)	—	—	—
	50 (4)	199 (8)	574 (5)
	537 (4)	223 (9)	831 (5)
H(3)	519 (12)	402 (10)	203 (7)
	-3 (13)	129 (10)	248 (7)
	95 (5)	464 (10)	432 (5)
	603 (5)	459 (10)	686 (6)
H(4)	531 (11)	212 (10)	200 (7)
	942 (12)	243 (10)	128 (7)
	229 (5)	479 (10)	439 (6)
	744 (6)	477 (10)	676 (5)
H(5)	145 (12)	216 (10)	65 (7)
	766 (11)	432 (10)	244 (7)
	317 (5)	246 (10)	667 (6)
	819 (5)	255 (10)	921 (6)
H(6)	128 (11)	401 (9)	67 (7)
	814 (11)	329 (9)	358 (7)
	189 (5)	199 (10)	673 (6)
	696 (5)	221 (10)	946 (6)
H(7)	441 (12)	523 (9)	435 (7)
	422 (12)	66 (10)	454 (8)
	439 (5)	- 58 (10)	565 (6)
	941 (5)	- 86 (10)	838 (6)
H(8)	727 (13)	16 (10)	217 (8)
	318 (15)	209 (10)	376 (9)
	389 (6)	-225 (11)	424 (6)
	902 (6)	-246 (11)	698 (6)

Table 3 (cont.)

H(9)	57 (14)	445 (12)	296 (8)
	483 (15)	330 (10)	371 (9)
	311 (6)	- 60 (12)	273 (6)
	880 (6)	- 74 (11)	534 (6)

symmetry-related molecules are bonded, *i.e.* no hydrogen bonds exist between the independent molecules IA and IB. This means that two intermeshed but independent systems of hydrogen bonded molecules exist in the crystal. These systems are the same except that N(1A) is bonded only to O(1A) while N(1B) is bonded to both O(1B) and O(2B). All the hydrogen bond distances have normal values. The hydrogen bonding network is shown in Fig. 4.

The crystal structure of, and the hydrogen bonding in, sulphathiazole III is very similar to that of sulphathiazole II. The nitrogen atom N(2) does not participate in any hydrogen bonding and its only close approach that can be regarded as a van der Waals interaction is the N(1A)···N(2B) distance of 3.223 Å. The independent molecules IIIA and IIIB are bonded into integrated, irregular sheets of molecules that are parallel to the (001) plane. The hydrogen bonding in sulphathiazole III is illustrated by the schematic diagram in Fig. 5. Neighbouring molecules are hydrogen bonded through N(1)···O(1) and N(1)···N(3) bonds. The absence of one N(1A)···O(1B) bond implies that the oxygen atom of molecule IIIA that participates in hydrogen bonding, is bonded to the amine nitrogen atoms of two separate IIIB molecules, while the oxygen atom of molecule IIIB is bonded to the amine nitrogen atom of only one IIIA molecule.

Table 4. Bond lengths (Å) and angles (°) with e.s.d.'s of sulphathiazole I and III

(a) Bond lengths

	IA	IB	IIIA	IIIB	Mean
N(1)-C(1)	1.368 (11)	1.382 (11)	1.399 (8)	1.422 (8)	1.394 (18)
C(1)-C(2)	1.390 (12)	1.419 (13)	1.386 (7)	1.385 (7)	1.395 (13)
C(2)-C(3)	1.365 (14)	1.363 (12)	1.398 (8)	1.370 (8)	1.375 (13)
C(3)-C(4)	1.386 (12)	1.389 (13)	1.393 (6)	1.399 (7)	1.392 (5)
C(4)-C(5)	1.384 (12)	1.377 (11)	1.393 (7)	1.386 (8)	1.385 (9)
C(5)-C(6)	1.375 (14)	1.355 (11)	1.350 (9)	1.395 (9)	1.369 (16)
C(6)-C(1)	1.390 (12)	1.395 (12)	1.430 (7)	1.375 (7)	1.397 (18)
C(4)-S(1)	1.749 (9)	1.748 (10)	1.745 (5)	1.769 (6)	1.754 (9)
S(1)-O(1)	1.432 (6)	1.436 (7)	1.450 (4)	1.440 (4)	1.440 (6)
S(1)-O(2)	1.437 (6)	1.415 (7)	1.446 (4)	1.435 (4)	1.434 (10)
S(1)-N(2)	1.615 (8)	1.612 (7)	1.610 (5)	1.602 (5)	1.606 (9)
N(2)-C(7)	1.316 (10)	1.296 (10)	1.321 (7)	1.316 (7)	1.312 (10)
S(2)-C(7)	1.733 (9)	1.736 (10)	1.727 (6)	1.752 (5)	1.738 (9)
C(7)-N(3)	1.349 (11)	1.369 (10)	1.344 (7)	1.343 (7)	1.347 (12)
N(3)-C(8)	1.357 (10)	1.371 (12)	1.385 (7)	1.371 (7)	1.372 (9)
C(8)-C(9)	1.297 (13)	1.324 (14)	1.327 (9)	1.327 (8)	1.318 (11)
C(9)-S(2)	1.742 (12)	1.705 (12)	1.735 (7)	1.737 (7)	1.728 (14)
N(1)-H(1)	—	—	0.82	0.90	0.87
N(1)-H(2)	—	—	0.98	0.96	0.93
C(2)-H(3)	1.00	1.04	0.91	0.98	0.97
C(3)-H(4)	1.06	0.98	0.88	1.28	1.03
C(5)-H(5)	0.91	0.90	0.90	0.78	0.88
C(6)-H(6)	1.03	0.94	1.02	1.04	0.99
N(3)-H(7)	1.06	0.93	0.89	1.03	0.96
C(8)-H(8)	1.00	0.87	1.00	1.00	0.96
C(9)-H(9)	0.89	0.79	1.20	1.40	1.03

Table 4 (cont.)

(b) Bond angles	IA	IB	IIIA	IIIB	Mean
N(1)-C(1)-C(2)	120.7 (9)	120.1 (9)	122.5 (5)	121.9 (5)	121.5 (9)
N(1)-C(1)-C(6)	121.9 (9)	122.6 (9)	119.4 (5)	117.2 (5)	120.1 (19)
C(1)-C(2)-C(3)	122.6 (9)	120.4 (9)	120.5 (5)	120.9 (5)	121.0 (8)
C(2)-C(3)-C(4)	119.3 (8)	120.8 (9)	120.1 (5)	119.1 (5)	119.9 (6)
C(3)-C(4)-C(5)	119.3 (9)	118.9 (9)	119.6 (5)	119.6 (5)	119.4 (4)
C(4)-C(5)-C(6)	120.8 (8)	121.2 (9)	120.5 (5)	121.0 (5)	120.7 (4)
C(5)-C(6)-C(1)	120.7 (9)	121.4 (9)	121.2 (6)	118.4 (5)	120.5 (11)
C(6)-C(1)-C(2)	117.4 (9)	117.3 (9)	118.0 (5)	120.9 (5)	118.4 (13)
C(3)-C(4)-S(1)	120.5 (7)	120.1 (7)	120.9 (4)	121.2 (4)	120.7 (4)
C(5)-C(4)-S(1)	120.2 (7)	120.8 (8)	119.5 (4)	119.1 (4)	119.9 (7)
C(4)-S(1)-O(1)	108.2 (4)	107.1 (4)	107.9 (2)	107.1 (3)	107.5 (5)
C(4)-S(1)-O(2)	107.3 (4)	108.1 (4)	107.6 (2)	108.3 (2)	107.8 (4)
O(1)-S(1)-O(2)	117.7 (4)	118.1 (5)	116.1 (2)	116.8 (3)	117.1 (8)
N(2)-S(1)-O(1)	104.2 (4)	103.1 (4)	105.7 (2)	105.8 (2)	104.8 (9)
N(2)-S(1)-O(2)	111.2 (4)	112.2 (4)	113.9 (3)	113.6 (3)	113.0 (11)
N(2)-S(1)-C(4)	107.8 (4)	107.7 (4)	105.0 (3)	104.3 (3)	105.9 (17)
S(1)-N(2)-C(7)	120.9 (5)	120.2 (5)	119.6 (4)	120.4 (4)	120.2 (5)
N(2)-C(7)-S(2)	130.3 (8)	132.0 (8)	131.1 (5)	129.8 (4)	130.8 (8)
N(2)-C(7)-N(3)	120.7 (10)	120.2 (9)	119.4 (5)	122.0 (5)	120.6 (8)
C(7)-N(3)-C(8)	115.3 (8)	114.6 (9)	115.1 (5)	116.7 (5)	115.6 (8)
N(3)-C(8)-C(9)	113.8 (10)	114.1 (10)	113.0 (6)	112.8 (6)	113.5 (5)
C(8)-C(9)-S(2)	111.7 (8)	111.0 (8)	111.1 (5)	111.4 (5)	111.3 (4)
C(9)-S(2)-C(7)	90.0 (5)	92.4 (6)	91.2 (3)	90.9 (3)	91.1 (8)
S(2)-C(7)-N(3)	109.0 (7)	107.8 (8)	109.5 (4)	108.1 (4)	108.6 (6)
H(1)-N(1)-C(1)	—	—	92	122	110
H(2)-N(1)-C(1)	—	—	101	118	112
H(3)-C(2)-C(1)	107	111	125	128	118
H(4)-C(3)-C(4)	120	117	112	111	116
H(5)-C(5)-C(4)	109	120	115	109	114
H(6)-C(6)-C(1)	122	115	122	126	121
H(7)-N(3)-C(7)	123	102	119	116	116
H(8)-C(8)-N(3)	138	134	124	117	127
H(9)-C(9)-C(8)	91	95	129	123	114

Comparison of these two crystal structures with that of sulphathiazole II shows that the main difference lies in the types of hydrogen bonds present and their effects on the arrangement of molecules in the crystal. Although sulphathiazole I has the additional N(2)···N(3) bond and uses O(2) in a N(1)···O(2) bond, the number of bonds per molecule is the same as in form III. The strengths of the hydrogen bonds as shown by their bond

distances are very similar and no definite conclusion as to changes in stability due to differing hydrogen bonding can be made. A similar situation has been encountered with the crystal structures of barbital I and II (Craven, Vizzini & Rodrigues, 1969) and amobarbital I and II (Craven & Vizzini, 1969). These structures are also characterized by the small differences between polymorphs.

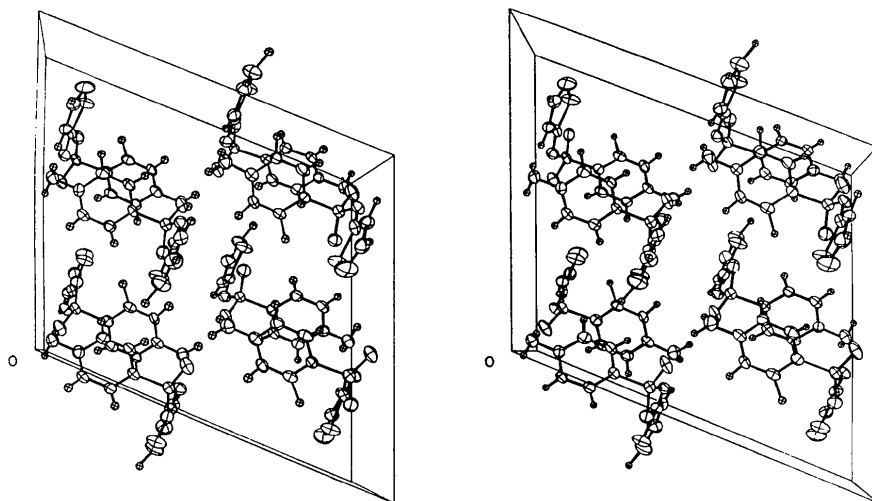


Fig. 3. A stereo pair showing the unit cell and its contents of sulphathiazole III viewed along the *b* axis with the *c* axis vertical.

Modification I is the most suitable form of sulphathiazole for pharmaceutical use owing to its higher rate of dissolution. It is clear from these structure analyses that this is not a result of striking differences in the numbers of hydrogen bonds in the different polymorphs but rather a consequence of more subtle bonding effects. Evidence of the weaker bonding in modification I lies in its density which is significantly (4%) lower than that of the other two polymorphs.

Table 5. *The least-squares planes through the benzene and thiazole rings for sulphathiazole I and III*

The equations of the planes are of the form $AX + BY + CZ + D = 0$ and were calculated with reference to the orthogonal coordinate system a, b, c^* . The deviations from the planes are in Å.

(a) Planes through the six benzene ring atoms

Coefficients	Molecules			
	IA	IB	III A	III B
A	-0.499	0.760	-0.125	-0.066
B	-0.070	0.528	0.830	0.869
C	0.863	0.379	0.544	0.491
D	0.151	9.525	6.793	7.849
C(1)	0.011	-0.009	-0.014	-0.012
C(2)	-0.006	0.018	0.010	0.005
C(3)	-0.007	-0.011	-0.003	0.003
C(4)	0.014	-0.005	0.001	-0.004
C(5)	-0.009	0.014	-0.006	-0.003
C(6)	-0.004	-0.007	0.013	0.011
N(1)	0.063	-0.041	-0.014	-0.003
S(1)	0.112	0.104	-0.017	0.033
H(3)	0.11	-0.10	0.00	-0.02
H(4)	0.17	0.01	-0.12	-0.13
H(5)	-0.04	0.16	-0.03	-0.05
H(6)	-0.08	0.06	-0.02	0.03

(b) Planes through the atoms C(7), N(3), C(8) and C(9)

Coefficients	Molecules			
	IA	IB	III A	III B
A	-0.347	-0.501	0.994	0.991
B	0.935	0.496	0.111	0.131
C	0.072	0.709	-0.001	-0.010
D	6.076	4.814	4.082	11.229
C(7)	0.002	-0.007	0.001	-0.002
N(3)	-0.004	0.013	-0.002	0.003
C(8)	0.004	-0.013	0.002	-0.003
C(9)	-0.002	0.007	-0.001	0.002
S(2)	0.102	0.037	0.013	0.037
N(2)	-0.054	-0.048	0.030	0.003
H(7)	0.08	-0.18	0.11	-0.10
H(8)	0.14	0.19	-0.06	-0.10
H(9)	0.11	0.05	-0.36	0.72

(c) Dihedral angles

IA	IB	III A	III B
101.26	98.62	91.87	100.04

Polymorphs II and III change into polymorph I on slow heating to 173–175°C. Rapid heating causes III to melt at this temperature with no apparent phase change before melting. The presence of the most active form (I) can thus be ensured by slow heating to temperatures between 175 and 200°C. Reversion from the metastable polymorph I to one of the other two forms takes

Table 6. *Intermolecular distances*

(a) Hydrogen bonding

IA		IB	
N(1A)···O(2A)	3.088 Å	N(1B)···O(1B)	2.940 Å
N(2A)···N(3A)	2.874	N(1B)···O(2B)	2.959
N(2A)···H(7A)	1.87	N(2B)···N(3B)	2.877
		N(2B)···H(7B)	1.95

III A		III B	
N(1A)···O(1B)	3.026 Å	N(1B)···O(1A)	3.010 Å
N(1A)···N(3B)	2.925	N(1B)···O(1A)	3.022
N(1A)···H(7B)	1.91	N(1B)···N(3A)	2.871
O(1A)···N(1B)	3.010	N(1B)···H(7A)	2.06
O(1A)···H(1B)	2.16	O(1B)···N(1A)	3.026
O(1A)···N(1B)	3.022	O(1B)···H(2A)	2.07
O(1A)···H(2B)	2.17	N(3B)···N(1A)	2.925
N(3A)···N(1B)	2.871		

(b) Other close approaches

Form I			
N(1A)···C(7B)	3.238 Å		
N(1A)···N(1B)	3.392	O(1B)···N(3B)	3.249
N(1A)···O(2B)	3.449		
N(1A)···N(2B)	3.451	C(4B)···C(7B)	3.383
O(1A)···C(6B)	3.369	C(5B)···N(2B)	3.316
O(1A)···C(5B)	3.375		
O(2A)···C(2A)	3.288		
C(4A)···C(7A)	3.389		
N(2A)···C(3A)	3.455		
N(3A)···C(9B)	3.285		

Form III			
N(1A)···N(2B)	3.223 Å	O(1B)···C(6A)	3.445
N(1A)···O(1B)	3.252	O(2B)···C(3A)	3.271
O(1A)···C(6B)	3.344	O(2B)···C(2A)	3.388
O(1A)···C(1B)	3.432	O(2B)···N(3B)	3.406
O(2A)···C(3B)	3.215	N(3B)···C(1A)	3.366
O(2A)···C(2B)	3.366	N(3B)···C(6A)	3.473
C(7A)···C(8A)	3.484		
N(3A)···N(3A)	3.266		
N(3A)···C(1B)	3.375		
N(3A)···C(6B)	3.429		
N(3A)···C(8A)	3.471		

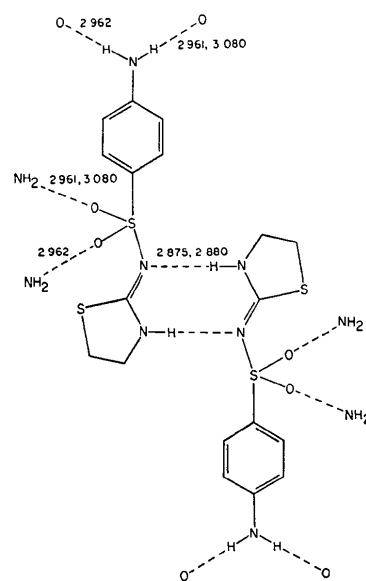


Fig. 4. The hydrogen bonding in sulphathiazole I.

place slowly at room temperature but the kinetics of the process has not been studied as extremely long times are involved.

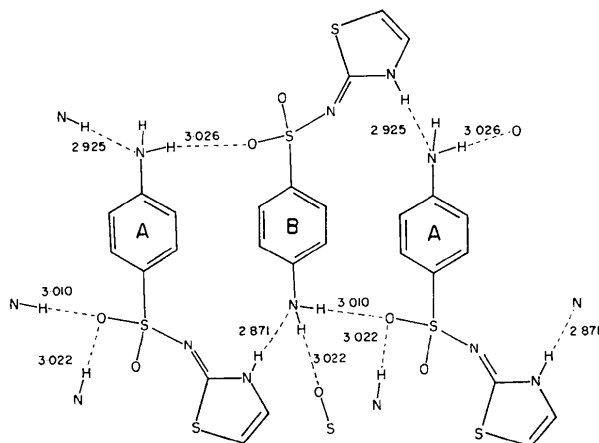


Fig. 5. The hydrogen bonding in sulphathiazole III.

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The Crystal and Molecular Structure of 2-Thioamidopyridine

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Crystals of 2-thioamidopyridine, $C_6H_6N_2S$, are monoclinic, space group $P2_1/c$ with four molecules in a unit cell of dimensions $a = 5.79$, $b = 7.42$ and $c = 16.26$ Å, $\beta = 106.9^\circ$. The crystal structure was solved by three-dimensional Patterson and Fourier syntheses using the heavy-atom technique. The atomic parameters were refined by block-matrix least-squares calculations. Hydrogen atoms were located but their parameters were not refined. The final R index was 0.069 from 1004 observed reflexions, collected using $Cu K\alpha$ radiation. The mean bond lengths are $C-C = 1.383$, $C-N = 1.336$ Å in the pyridine ring and $C-C = 1.505$, $C-N = 1.325$, $C-S = 1.657$ Å in the thioamido group with a mean e.s.d. of 0.006 Å. Bond lengths and angles in the thioamido group suggest that the electron configuration contains contributions from the ionic canonical form as well as the amide form. The angle between the plane of the pyridine ring and that of the thioamido group is 10.5° . Non-bonded contact radii of some atoms are proposed to interpret the twisting between the thioamido group and the pyridine ring. Pairs of hydrogen bonds of the type $N-H \cdots S$ and 3.43 Å in length join molecules together across centres of symmetry.

Introduction

The corrosion inhibitive properties of aliphatic and aromatic amines, as well as nitrogen heterocyclic compounds, have been known for some time. (Rhodes & Khun, 1929; Mann, Lauer & Hultin, 1936). Sulphur-containing compounds such as thiourea derivatives

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and mercapto-imidazole derivatives have also been extensively used as inhibitors. (Aronson & Belen'kii, 1964). In view of the chelating ability of 2-thioamidopyridine (2-thiopic) to group IIB metals (Sutton, 1963) and certain transition metals (Sutton, 1966) the corrosion inhibitive properties of 2-thiopic were investigated, the compound was found to reduce the corrosion rate of carbon steels in 0.2*N* sulphuric acid by 40%. (Osborne, 1968).

In its chelation with metals, 2-thiopic always uses its pyridine nitrogen atom as well as either the nitrogen atom or the sulphur atom of the side chain. Spectro-

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